

Formaldehyde production from isoprene oxidation across NO_x regimes

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Abstract

The chemical link between isoprene and formaldehyde (HCHO) is a strong, non-linear function of NO_x ($= \text{NO} + \text{NO}_2$). This relationship is a linchpin for top-down isoprene emission inventory verification from orbital HCHO column observations. It is also a benchmark for overall photochemical mechanism performance with regard to VOC oxidation. Using a comprehensive suite of airborne *in situ* observations over the Southeast U.S., we quantify HCHO production across the urban-rural spectrum. Analysis of isoprene and its major first-generation oxidation products allows us to define both a “prompt” yield of HCHO (molecules of HCHO produced per molecule of freshly-emitted isoprene) and the background HCHO mixing ratio (from oxidation of longer-lived hydrocarbons). Over the range of observed NO_x values (roughly 0.1 – 2 ppbv), the prompt yield increases by a factor of 3 (from 0.3 to 0.9 ppbv ppbv⁻¹), while background HCHO increases by a factor of 2 (from 1.6 to 3.3 ppbv). We apply the same method to evaluate the performance of both a global chemical transport model (AM3) and a measurement-constrained 0-D steady state box model. Both models reproduce the NO_x dependence of the prompt HCHO yield, illustrating that models with updated isoprene oxidation mechanisms can adequately capture the link between HCHO and recent isoprene emissions. On the other hand, both models under-estimate background HCHO mixing ratios, suggesting missing HCHO precursors, inadequate representation of later-generation isoprene degradation and/or under-estimated hydroxyl radical concentrations. Detailed process rates from the box model simulation demonstrate a 3-fold increase in HCHO production across the range of observed NO_x values, driven by a 100% increase in OH and a 40% increase in branching of organic peroxy radical reactions to produce HCHO.

1 Introduction

Formaldehyde (HCHO) is a ubiquitous byproduct of volatile organic compound (VOC) oxidation. While methane is the principal HCHO precursor in remote regions, larger VOC are the main source over continents. HCHO is also directly emitted via biomass burning (Lee et al., 1997), fossil fuel combustion (Luecken et al., 2012), natural gas flaring (Knighton et al., 2012), ethanol refining (de Gouw et al., 2015), possibly vegetation (DiGangi et al., 2011) and agricultural activity (Kaiser et al., 2015a), but chemical production dominates the global budget (Fortems-Cheiney et al., 2012). Photolysis and reaction with OH destroy HCHO with

1 a characteristic lifetime of several hours during midday, implying that the HCHO abundance
2 reflects recent VOC oxidation.

3 Globally, isoprene is the main precursor of near-surface HCHO. A highly reactive
4 diene emitted by vegetation, isoprene comprises roughly one third of all non-methane VOC
5 emissions (Guenther et al., 2012). Oxidation of isoprene in the presence of nitrogen oxides
6 ($\text{NO}_x = \text{NO} + \text{NO}_2$) stimulates the production of ozone (Trainer et al., 1987) and organic
7 aerosol precursors (Xu et al., 2015), impacting air quality and climate in many continental
8 regions. Biogenic emission inventories struggle to accurately represent the spatiotemporal
9 variability of isoprene emissions, with model-measurement discrepancies and differences
10 among emission inventories approaching a factor of 2 or more (Carlton and Baker, 2011;
11 Warneke et al., 2010). Such differences directly impact predicted ozone and aerosol
12 distributions (Hogrefe et al., 2011).

13 Numerous studies have applied satellite-based HCHO column observations as a top-
14 down constraint on isoprene emissions (see Kefauver et al. (2014) for a review). Typically, a
15 chemical transport model is employed both to supply *a priori* HCHO vertical distributions for
16 satellite retrievals (González Abad et al., 2015) and to relate HCHO column concentrations to
17 isoprene emission strength. Early studies utilized linear steady-state relationships (Palmer et
18 al., 2003), while recent computational advances have permitted full inversions that more fully
19 account for transport, multiple sources and varying chemical regimes (Fortems-Cheiney et al.,
20 2012). Such techniques have informed isoprene emission inventories in North America
21 (Abbot et al., 2003; Millet et al., 2008; Millet et al., 2006; Palmer et al., 2006; Palmer et al.,
22 2003), South America (Barkley et al., 2013; Barkley et al., 2008), Europe (Curci et al., 2010;
23 Dufour et al., 2009), Africa (Marais et al., 2012), Asia (Fu et al., 2007; Stavrou et al.,
24 2014), and globally (Fortems-Cheiney et al., 2012; Shim et al., 2005; Stavrou et al., 2009).
25 Future geostationary observations, such as the NASA Tropospheric Emissions: Monitoring of
26 Pollution (TEMPO, <http://science.nasa.gov/missions/tempo/>) mission, will permit an even
27 more detailed investigation of the spatial and temporal variability of isoprene emissions and
28 other VOC sources.

29 Chemistry dictates the relationship between HCHO columns and underlying isoprene
30 emissions. Many of the above-listed studies apply 0-D box model calculations to evaluate the
31 yield of HCHO from isoprene as a function of oxidation time, NO_x regime and chemical
32 mechanism. In all cases, it is found that NO_x enhances both the production rate and ultimate

yield of HCHO. Slower production at lower NO_x can lead to “smearing,” whereby HCHO production is displaced relative to the isoprene source. Palmer et al. (2003) define a characteristic smearing length scale, which can range from 10 to 100 km or more. Furthermore, accumulation of oxygenated VOC over multiple generations of isoprene degradation can contribute to substantial background HCHO production, which is not directly linked with fresh isoprene emissions. Long-lived primary anthropogenic or biogenic emissions, like methane and methanol, can also contribute to this background. Background column concentrations are typically on the order of $5 \times 10^{15} \text{ cm}^{-2}$, equal to 20% or more of the isoprene-driven HCHO column enhancement (Barkley et al., 2013; Millet et al., 2006). A wave of recent theoretical (Peeters et al., 2014; Peeters and Müller, 2010; Peeters et al., 1999), laboratory (Crounse et al., 2012; Crounse et al., 2011; Paulot et al., 2009a; Paulot et al., 2009b) and field (Mao et al., 2012) research has highlighted shortcomings in low-NO_x isoprene oxidation schemes. Such issues translate directly into top-down emission estimates; for example, Marais et al. (2012) report an uncertainty of 40% in satellite-derived African isoprene emissions at high-NO_x and 40-90% at low-NO_x. Coarse resolution of averaged satellite observations and model simulations (typically $1^\circ \times 1^\circ$ or more) has partly mitigated these problems in prior work, as variability in NO_x-dependent smearing and background production is averaged out. A more careful treatment will be needed to harness the enhanced resolution of near-future orbital observations (e.g., $8 \times 4.5 \text{ km}^2$ for TEMPO), especially since these measurements will include diurnal variability.

Here, we use a comprehensive set of *in situ* observations to quantify the impact of NO_x on the isoprene-HCHO chemical link. Using isoprene and its unique first-generation products, we segregate HCHO into two categories. The first, defined as “prompt” HCHO, is produced from fresh isoprene emissions (on a timescale of less than a day) and retains the signature of isoprene emission source strength. The second category is “background” HCHO stemming from oxidation of longer-lived isoprene oxidation products and other VOC. We examine the NO_x dependence of both quantities. Applying the same method to 0-D and global model simulations, we evaluate the ability of current chemical mechanisms to replicate the observed trends. Box model results are also used to elucidate the mechanistic underpinnings of the NO_x influence on HCHO production.

2 SENEX Observations

The Southeast Nexus (SENEX) mission was an airborne campaign designed to examine the interaction of natural and anthropogenic emissions (Warneke et al., 2016). During June and July of 2013, the NOAA WP-3D aircraft logged about 120 flight hours over 20 research flights in a range of environments throughout the Southeast United States, including urban centers, power plant plumes, natural gas extraction regions, agricultural areas and forests. The payload included a suite of gas- and particle-phase instrumentation. Here we utilize observations of HCHO, isoprene, methyl vinyl ketone (MVK), methacrolein (MACR), NO and NO₂. HCHO was measured at 1 Hz by the NASA In Situ Airborne Formaldehyde (ISAF) instrument, which utilizes the laser-induced fluorescence technique and has an accuracy of $\pm 10\%$ (Cazorla et al., 2015). Isoprene, MVK and MACR were measured by both a quadrupole proton transfer reaction mass spectrometer (PTR-MS) and the NOAA improved whole-air sampler (iWAS) with offline gas chromatography. The PTR-MS (de Gouw and Warneke, 2007) has a stated accuracy of 20% and sequentially sampled masses for isoprene ($m/z + 69$) and the sum of MVK and MACR ($m/z + 71$) for 1 s each with a duty cycle of 14 s. The iWAS (Lerner et al., 2016) collected 72 canister samples each flight, which were analyzed offline with gas chromatography – mass spectrometry 3-4 days post-flight. iWAS measurement uncertainty is 20% for speciated MVK and MACR and 27% for isoprene. NO and NO₂ were measured at 1 Hz via chemiluminescence coupled with a photolytic NO₂ converter (Pollack et al., 2010; Ryerson et al., 1999) with an accuracy of 5%. Data are filtered to include only daytime boundary layer conditions (solar zenith angle $< 60^\circ$, radar altitude < 1 km). Influence from biomass burning (acetonitrile > 210 pptv and CO > 300 ppbv) is also removed. This procedure, along with the disjunct nature of the PTR-MS measurement, excludes 50% of all fast (1 Hz) data. After accounting for data gaps, we retain 8435 1 Hz data points and 81 iWAS samples.

Measurements of MVK and MACR may include a positive bias from conversion of isoprene hydroxyhydroperoxides (ISOPOOH) on hot metal surfaces in the sampling system (Liu et al., 2013; Rivera-Rios et al., 2014). ISOPOOH mixing ratios up to 2 ppbv were observed by the University of Washington Iodide high-resolution time-of-flight chemical ionization mass spectrometer during SENEX. Neither the NOAA PTR-MS nor the iWAS have been tested for this interference with an ISOPOOH standard, thus we cannot definitively rule out such artifacts or develop a correction factor. To our knowledge, it is not yet clear how

1 the putative interference depends on instrument configuration (flow rates, electric fields, etc.).
2 Thus, caution is warranted when comparing the SENEX systems to similar, but not identical,
3 instruments. Theoretically, this mechanism could give rise to an analogous artifact in HCHO
4 observations. Recent laboratory tests, however, indicate that the ISOPOOH-to-HCHO
5 conversion efficiency in ISAF is less than 5% (St. Clair et al., 2016).

6 We cannot unambiguously quantify the ISOPOOH artifact using observations alone,
7 but we can gain some insight from comparing PTR-MS and iWAS data. On average, iWAS
8 observations of MVK+MACR are ~40% higher than those from the PTR-MS (Figs. S1 and
9 S2), suggesting a systematic bias in one or both measurements. Both instruments were
10 calibrated using the same gas standards, and the two techniques agree well for other species
11 such as isoprene (Lerner et al., 2016; Warneke et al., 2016), so a calibration error is unlikely.
12 Production of oxygenated VOC in ambient air samples collected and aged in stainless steel
13 canisters cannot be ruled out. Enhancements in MVK and MACR (above the 20%
14 uncertainty) have been observed in canisters after aging over ~11 days (Lerner et al., 2016),
15 though this is significantly longer than typical turn-around times for SENEX. To evaluate the
16 potential for ISOPOOH conversion to explain this discrepancy, we plot the ratio and
17 difference of the PTR-MS and iWAS measurements as a function of ISOPOOH in Fig. S2.
18 While the ratio is essentially constant (iWAS/PTR-MS ~1.43), the absolute difference
19 exhibits a strong positive correlation with ISOPOOH ($r^2 = 0.43$). The slope of this relationship
20 implies that a conversion of 50% of ISOPOOH to MVK and/or MACR in the iWAS system
21 would explain the difference in the two measurements. Correcting total iWAS MVK+MACR
22 for such an artifact reduces the slope of the iWAS-PTR-MS correlation from 1.48 to 1.24
23 (Fig. S1B), bringing agreement to well with combined measurement uncertainties. In practice,
24 we cannot apply such a correction to the speciated iWAS observations as the conversion
25 efficiency may be different for each isomer. This result does not exclude the possibility of an
26 artifact in the PTR-MS measurement, though it does suggest an upper limit ISOPOOH
27 conversion efficiency of 50% for the PTR-MS (which would imply a conversion of 100% for
28 the iWAS). The analysis presented in Sections 3 and 4 primarily relies on PTR-MS data due
29 to its greater temporal coverage. Our key conclusions are not impacted by a 50% ISOPOOH
30 correction to the PTR-MS data, thus we use the data without correction.

31 SENEX sampled a wide spectrum of chemical regimes (Figure 1). For the daytime
32 boundary-layer observations presented here, maximum 1 Hz isoprene and NO mixing ratios

1 respectively reach 8.1 and 95 ppbv, while minima are less than a few pptv. The distributions
2 of both isoprene and NO observations are approximately log-normal (top and right panels of
3 Fig. 1), peaking at 1.5 ppbv and 50 pptv, respectively. Though these distributions may be
4 biased towards areas of urban influence, the range of environments encountered during
5 SENEX is representative of the Southeast U.S. summertime boundary layer. The long tail at
6 the low end of the isoprene distribution is mostly associated with regions lacking significant
7 tree cover where isoprene emissions are lower, notably Illinois and Indiana. The NO
8 distribution spans four orders of magnitude (< 10 to $\sim 10^4$ pptv), over which radical chemistry
9 changes markedly. At NO mixing ratios of a few hundred pptv or more, organic peroxy
10 radicals (RO_2) react mostly with NO. At low NO (10's of pptv or less), reaction with HO_2 ,
11 other RO_2 and isomerization dominate the RO_2 fate. The bulk of the NO distribution lies in a
12 transition region for radical chemistry, making this dataset ideal for probing the
13 anthropogenic influence on biogenic VOC oxidation.

14 HCHO mixing ratios (color shading in Fig. 1) range from 0.8 to 14 ppbv with a mean
15 of 4.3 ppbv. HCHO is most abundant in regions where both isoprene and NO_x are elevated.
16 High NO_x is often accompanied by increased concentrations of anthropogenic VOC; however,
17 constrained box-model calculations demonstrate that isoprene is the dominant HCHO
18 precursor even in these cases (Sect. 5). Thus, changes in radical cycling and partitioning (and
19 not co-variance of NO_x and anthropogenic VOC) drives the observed NO_x dependence of
20 HCHO abundance.

22 3 Linking Observed and Emitted Isoprene

23 The isoprene photochemical cascade is a multi-step process. Isoprene oxidation is initiated via
24 reaction with the hydroxyl radical (OH), ozone, or the nitrate radical (NO_3). In the Southeast
25 U.S., typical daytime levels for OH, ozone and NO_3 are $4 \times 10^6 \text{ cm}^{-3}$, 50 ppbv and 0.1 pptv,
26 respectively (OH and NO_3 are estimated from median box model output, see Sect. 5). The
27 corresponding isoprene lifetimes at 298K are 0.7 h, 17 h and 160 h, respectively. Thus,
28 reaction with OH typically constitutes 95% or more of the total daytime isoprene sink in this
29 environment. Addition of OH and reaction with O_2 generates one of several isoprene
30 hydroperoxy radicals (ISOPO_2). ISOPO_2 isomers interconvert rapidly due to reversible O_2
31 addition (Peeters et al., 2009) but are eventually destroyed via reaction with NO, hydroperoxy
32 radical (HO_2), other organic peroxy radicals (RO_2) or isomerization. Most branches have the

potential to produce HCHO, with varying yields. The laboratory-derived first-generation HCHO yield from the NO pathway is ~0.6 (Atkinson and Arey, 2003), though this value may be less representative of the real atmosphere due to the very high isoprene concentrations (and very short RO₂ lifetimes) in early chamber experiments. The first-generation yield from the HO₂ pathway is ~0.06 (Liu et al., 2013). Isomerization chemistry is less well understood; the 1,5-H-shift is believed to produce HCHO with a unity yield, while the much faster 1,6-H-shift should not produce any HCHO (da Silva et al., 2010; Fuchs et al., 2013; Peeters et al., 2014; Peeters and Müller, 2010; Peeters et al., 2009). Regardless of the specific pathway, MVK or MACR are always co-produced with HCHO in the first generation. HCHO is also generated in subsequent chemistry, but on a longer timescale and from a much larger suite of precursors. For example, the OH lifetimes of MACR and MVK are respectively 3.5 and 5 times longer than that of isoprene. HCHO, MVK and MACR are also high-yield products of isoprene ozonolysis (Atkinson and Arey, 2003), but as noted above this reaction is relatively slow. Nighttime oxidation of isoprene by NO₃ radical is also likely a negligible source of these carbonyls (Brown et al., 2009). Yields are small (Atkinson and Arey, 2003; Kwok et al., 1996), and the lifetimes of MVK, MACR and HCHO are sufficiently short that any nighttime production should not influence the midday observations considered here.

Boundary layer composition reflects a mixture of emissions with various degrees of photochemical processing. To isolate the impact of “fresh” isoprene emissions, we exploit the relatively simple chemistry of MVK and MACR, which are produced via isoprene (ISOP) oxidation and lost primarily via reaction with OH.



Rate constants (k) are taken from the IUPAC database (Atkinson et al., 2006). These reactions form the basis for a photochemical clock of isoprene oxidation (de Gouw et al., 2005; Roberts et al., 2006; Stroud et al., 2001). Integration of the kinetic equations for this system shows that the product/parent ratios are a function of the rate constants, yield (y), reaction time (t) and mean OH concentration. In the case of MACR, for example:

$$\frac{[\text{MACR}]}{[\text{ISOP}]} = \frac{y_{\text{MACR}}k_1}{k_2 - k_1} (1 - \exp((k_1 - k_2)[\text{OH}]t)) \quad (1)$$

1 An analogous expression holds for MVK. As noted by Stroud et al. (2001), this “sequential
2 reaction model” is purely chemical and does not account for the effects of mixing and
3 transport. Indeed, this analysis relates daughter/parent ratios to an “average” photochemical
4 age, when in fact there is a broad distribution of ages in any mixed air mass. We also
5 implicitly assume that direct emissions (Fares et al., 2015) and deposition (Karl et al., 2010)
6 of MVK and MACR do not significantly influence the budget of these compounds.

7 Two potential issues arise when applying this model to the real atmosphere. First, the
8 yields of MVK and MACR are dependent on ISOPO₂ branching and are thus a non-linear
9 function of NO_x. Previous applications of this method (de Gouw et al., 2005; Roberts et al.,
10 2006; Stroud et al., 2001) have assumed lab-derived high-NO_x yields of 0.33 and 0.23 for
11 MVK and MACR, respectively (Atkinson and Arey, 2003), but this may not be appropriate in
12 the present case; furthermore, these yields are not fully consistent with current chemical
13 mechanisms. Given the wide range of conditions sampled, we explicitly account for NO_x-
14 dependent yields for MVK and MACR. For this purpose, we conducted a series of pseudo-
15 chamber simulations using a box model driven by the Master Chemical Mechanism (MCM)
16 v3.3.1 (Jenkin et al., 2015). As described in the SI, model setup mimics typical daytime
17 conditions in the Southeast U.S. (Fig. S3B), and yields are derived using a standard
18 procedure. Resulting yield curves (Fig. S3A) are then interpolated to observed NO mixing
19 ratios. Second, the photochemical age (*t*) implied by any observed daughter/parent ratio
20 depends on the concentration of OH, which was not measured and varies as an air mass ages.
21 Rather than assume a single “typical” value for OH, we express photochemical age in terms
22 of “exposure,” defined here as the product of OH concentration and reaction time averaged
23 over the photochemical lifetime of an air mass.

24 Figure 2 compares the observed relationship of iWAS MVK/isoprene and
25 MACR/isoprene ratios against theoretical trends predicted by the sequential reaction model.
26 Theoretical ratios are calculated at fixed exposures of 2, 4, 8, 12 and 16 × 10⁶ OH cm⁻³ h
27 using model derived yields for the 5th/95th percentiles of the observed NO distribution (NO =
28 20/200 pptv, *y*_{MVK} = 0.18/0.38, *y*_{MACR} = 0.11/0.20). Observed ratios of MVK/isoprene versus
29 MACR/isoprene exhibit a tight linear correlation. Higher ratios are often associated with
30 higher NO_x, likely reflecting enhanced OH and higher product yields in these air masses. Far
31 downwind from isoprene and NO_x source regions, we would expect to see higher
32 MVK/isoprene and MACR/isoprene ratios associated with lower NO_x due to removal of the

latter. The theoretical slope agrees well with observations, indicating exposures of $1 - 16 \times 10^6$ OH cm⁻³ h. For a typical daytime OH concentration of 4×10^6 cm⁻³, this corresponds to processing times of 0.25 – 4 hours.

The assumed MVK and MACR yields dictate the correspondence between daughter/parent ratios and exposure. For example, a MACR/isoprene ratio of 1 would be consistent with an exposure of 7.9×10^6 OH cm⁻³ h at NO = 20 pptv versus 6.0×10^6 OH cm⁻³ h at NO = 200 pptv. Thus, for any given daughter/parent ratio, a higher assumed yield gives a smaller derived exposure. The ratio of y_{MVK} to y_{MACR} determines the location of the theoretical line, and the excellent agreement of this relationship with observations in Fig. 2 indicates that MCMv3.3.1 accurately represents the branching ratios for MVK and MACR production within the sampled NO_x range.

We can effectively reverse this photochemical clock to derive a proxy for the total isoprene emissions that have been released into the sample air masses (de Gouw et al., 2005). First, we calculate OH exposures from observed daughter/parent ratios by inverting Eqn. (1). To perform this calculation with PTR-MS data (which has far greater coverage than the iWAS), we partition the measured sum between MVK and MACR using MVK/MACR ratios from steady-state box model calculations (Sect. 5). Modeled MVK/MACR ratios (with an output interval of 1 minute) are linearly interpolated to the 14-second observational time base. The MVK/MACR ratio does not vary dramatically (mean $\pm 1\sigma$: 2.3 ± 0.2), and using a constant ratio instead alters exposures by less than 4%. Calculated exposures range from 1 to 20×10^6 OH cm⁻³ h (Fig. S4A). Exposures derived from MACR are 6% lower than those from MVK on average, and we use the mean of these two values. Next, an “initial” isoprene mixing ratio, $ISOP_0$, is estimated via reverse integration of isoprene’s first-order loss rate:

$$[ISOP]_0 = [ISOP] \exp(k_1 [OH] t) \quad (2)$$

$ISOP_0$ represents the amount of isoprene that an air parcel would have to start with to generate the amount of isoprene, MVK and MACR observed. Thus, it is an observationally-constrained surrogate for isoprene emission strength (modulated to some degree by boundary layer height, as it is a volume-based quantity). $ISOP_0$ mixing ratios are typically 2 – 20 times higher than observed isoprene (Fig. S4B).

4 The Yield of HCHO from Isoprene

The definition of “yield” can vary with context and requires careful consideration when quantifying chemical relationships. In a mechanistic sense, the “first generation yield” refers to the amount of HCHO produced per unit isoprene consumed in the first stage of oxidation. This is analogous to the yields of MVK and MACR used in the above calculation of initial isoprene. The model-derived first-generation HCHO yield from isoprene varies by more than a factor of 2 over the range of chemical environments encountered during SENEX (Fig. S3). An alternative definition is that of the “total yield” (sometimes referred to as the “molar yield,” e.g. Millet et al. (2006)), a time-dependent quantity that describes the total amount of HCHO produced over multiple generations of oxidation. The total yield is typically derived from model simulations and used to relate satellite HCHO column observations to isoprene emissions (Marais et al., 2012; Millet et al., 2006). Early studies acknowledged the NO_x dependence of the total yield (Millet et al., 2006; Palmer et al., 2003), and more recent work has attempted to account for this dependence using NO_2 column observations (Marais et al., 2012). Here, we define the “prompt yield” as the change in observed HCHO per unit change in ISOP_0 ($\Delta\text{HCHO}/\Delta\text{ISOP}_0$). This is not the same as the first-generation yield, since the prompt yield can include HCHO production and loss over several hours (depending on the photochemical exposure of an air mass). Nor is it the same as the total yield, which inherently does not account for HCHO loss as an air mass ages. The prompt yield is effectively a quantity that relates isoprene emission strength to observed HCHO abundance. As we will demonstrate, this quantity is well-suited for segregating the various drivers of HCHO and for benchmarking model performance.

Figure 3A shows the relationship between calculated ISOP_0 and observed HCHO. The overall correlation is linear with a striking NO_x gradient. To quantify this NO_x dependence, we sort the data by $\log(\text{NO}_x)$, group it into 20 bins such that each bin contains the same number of points ($N = 416$), and perform a major-axis linear fit of HCHO versus ISOP_0 for each bin. Individual fits give r^2 values of 0.6-0.8, except for the highest NO_x bin ($r^2 = 0.48$) that contains some heavily-polluted air masses, such as downwind from power plants. Very fresh power plant plumes, defined as $\log(\text{NO}_x)$ values exceeding a mean + 3σ threshold, are removed prior to this procedure to avoid skewing the highest NO_x bin. Results are independent of the number of bins chosen or time resolution (e.g., 1-second versus 1-minute data).

1 The HCHO-ISOP₀ slope (Fig. 3B) represents the prompt yield. This yield varies by a
2 factor of 3 over the range of observed NO_x, from 0.3 ppbv ppbv⁻¹ for NO_x mixing ratios of a
3 few hundred pptv to 0.9 ppbv ppbv⁻¹ at NO_x > 1 ppbv. At low NO_x, the prompt yield is
4 comparable to the MCM-predicted direct first-generation yield of HCHO (0.25-0.4 ppbv
5 ppbv⁻¹ at NO = 10-40 pptv, Fig. S3), while at high NO_x it is somewhat higher than the
6 predicted first-generation yield (0.75 ppbv ppbv⁻¹ at NO = 1000 pptv). This likely reflects the
7 inclusion of more than one generation of HCHO production at higher NO_x, where oxidation is
8 more rapid (median exposures increase by 38% over the range of observed NO_x values). Most
9 of this portion of the HCHO budget, however, stems from first-generation production.

10 The intercept (Fig. 3C) represents the abundance of “background” HCHO. This
11 portion of the HCHO budget stems mainly from air that either has not encountered strong
12 isoprene emissions or is so aged that most of the isoprene has reacted away and can no longer
13 be linked to a specific source region. Some of this background may also stem from oxidation
14 of long-lived primary emissions like methane or methanol. Box model calculations (Sect. 5)
15 indicate average HCHO budget contributions of 0.3 ± 0.2 ppbv and 0.2 ± 0.1 ppbv from
16 methane and methanol, respectively. Background HCHO also exhibits a marked NO_x
17 dependence, increasing from 1.6 to 3.3 ppbv over the observed NO_x range. As with the
18 prompt yield, we expect such behavior since NO_x regulates the fate of all organic peroxy
19 radicals (see Sect. 6). Assuming a 1 km mixed layer depth (Wagner et al., 2015), the
20 corresponding HCHO column density for this background is $4 - 8 \times 10^{15} \text{ cm}^{-2}$. This is
21 comparable to the background reported by previous investigations of satellite-derived HCHO
22 columns (Barkley et al., 2013; Millet et al., 2006). None of these studies explicitly account for
23 the NO_x dependence of the background, though it can represent a substantial fraction of the
24 total HCHO column – maximum summertime HCHO columns over the southeast U.S. are
25 $\sim 25 \times 10^{15} \text{ cm}^{-2}$ (Millet et al., 2008). Given the strong NO_x dependence of both prompt and
26 background HCHO, grouping HCHO column observations by NO_x (e.g. using simultaneous
27 observations of NO₂ columns (Marais et al., 2012) or model-derived NO_x) and performing an
28 analysis similar to that described here could provide a robust means of accounting for these
29 influences.

5 Model Evaluation

Next, we compare the observed HCHO-ISOP₀ relationship to results from a global chemical-transport model and a 0-D box model. Our goals are to both illustrate the utility of this analysis and evaluate model performance. By going beyond a simple comparison of modeled and measured mixing ratios, we can more accurately pinpoint potential shortcomings in model chemistry.

The GFDL AM3 model is an atmospheric general circulation model with interactive chemistry (Donner et al., 2011), including recent updates to the representation of isoprene degradation (Mao et al., 2013; Naik et al., 2013). Model simulations were carried out at $50 \times 50 \text{ km}^2$ resolution with horizontal winds nudged to NCEP GFS analyses and sampled along the SENEX flight tracks at a time resolution of 1 minute. Further details are available elsewhere (Li et al., 2016).

The University of Washington Chemical Box Model (UWCM v2.2) is a versatile 0-dimensional framework for simulating various chemical systems, including lab chamber experiments (Wolfe et al., 2012) and observations from ground (Kim et al., 2015; Kim et al., 2013; Wolfe et al., 2014) and airborne (Marvin et al., 2015) platforms. Multiple chemical mechanisms are available within UWCM; here we used the latest version of the Master Chemical Mechanism (MCM v3.3.1, Jenkin et al. (2015)). UWCM was constrained with 1-minute average observations of isoprene, NO₂, ozone, CO, PAN, methane, methanol and meteorology and assumed clear-sky conditions for photolysis frequencies. The chemical system was integrated forward in time to diel steady state (total integration time of 3 days) for each set of measurements. This setup inherently assumes that the atmosphere is in chemical steady state – that is, that production and loss of HCHO, MVK, MACR and other species are roughly balanced. This assumption is rarely strictly true and may fail for highly-aged air masses (where isoprene is depleted) or when close to strong local emissions. Nonetheless, it is a fair approximation for the daytime well-mixed boundary layer observations that prevailed during SENEX. Monoterpenes and anthropogenic VOC are excluded from the simulation since observations of these species (from the iWAS) are relatively sparse. Separate sensitivity simulations utilizing the iWAS data suggest that observed monoterpenes and anthropogenic VOC (a subset of alkanes, alkenes and aromatics) increase modeled HCHO by $1 \pm 2\%$ and $2 \pm 3\%$, respectively. A more detailed evaluation of box model performance is forthcoming (Marvin et al., 2015).

Output from both models is filtered for daytime, boundary-layer, non-biomass burning points using the same criteria as that for observations (Sect. 2). Both models adequately reproduce observed HCHO mixing ratios (Fig. S5). We perform the same analyses as described above to derive model prompt yield and background HCHO. Because of the reduced time resolution, we group results into 10 NO_x bins, instead of 20, before fitting. For AM3, this results in 172 points per bin and typical r^2 values of 0.4 – 0.8. For UWCM, there are 134 points per bin and all r^2 values are > 0.86.

Both AM3 and UWCM reproduce the observed NO_x dependence of the prompt yield (Fig. 4A). AM3 agrees well with observations in both magnitude and trend, though with some scatter at mid-NO_x levels. UWCM tends to be slightly low throughout most of the NO_x range, which may reflect an issue with the mechanism (discussed below) or an inherent shortcoming of the steady-state assumption. Regardless of minor differences, these results suggest that both models provide excellent representation of early generation isoprene oxidation across NO_x regimes.

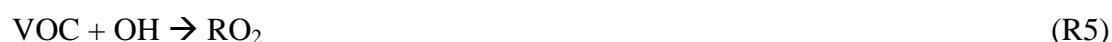
Background HCHO mixing ratios are under-predicted by 0.5 – 1 ppbv by both models (Fig. 4B). The range of under-prediction is consistent with the offsets between observed and modeled total HCHO abundances (Fig. S5 fit x-intercepts: 0.3 ppbv (AM3) and 1.1 ppbv (UWCM)). It is possible that both models are missing some HCHO precursors (e.g. from multi-generation isoprene oxidation or other VOC not related to isoprene). This is especially plausible for the UWCM simulation, which only includes isoprene, methane and methanol as primary VOC and does not account for horizontal transport. Under-estimated OH concentrations might also explain part of this discrepancy, though we cannot easily evaluate this possibility. AM3 performs somewhat better than UWCM in terms of overall magnitude but exhibits a less clear NO_x trend, which may reflect dilution over fairly large grid scales (note that the range of binned NO_x values is smaller for AM3 than both observations and the UWCM). This result again highlights the need to consider this background before using a model to interpret observed HCHO columns that effectively average HCHO sources over space and time.

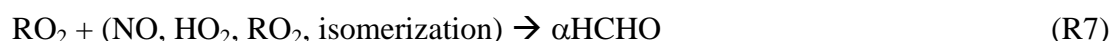
The agreement between AM3 and UWCM-MCMv3.3.1 is consistent with how these mechanisms treat first-generation ISOPO₂ radicals (Figs. S6 and S7). Both models use the same rate constants for reactions of ISOPO₂ with NO and HO₂, which comprise the bulk of ISOPO₂ sink. The AM3 mechanism assigns a 12% yield of HCHO to the reaction of ISOPO₂

with HO₂ (Paulot et al., 2009b), while the MCM assumes 100% production of peroxides for this channel. This may explain some of the discrepancy in the prompt yield at low NO_x (Fig. 4A), though neither mechanism is consistent with the current experimental HCHO yield of ~6% HCHO (Liu et al., 2013). There are also two key differences in the minor reaction channels. First, the rate constant for reaction of ISOPO₂ with other RO₂ is an order of magnitude lower in AM3 compared to MCMv3.3.1 (1.54 vs. 12 – 16 × 10⁻¹³ cm³ s⁻¹, the latter depending on the ISOPO₂ isomer distribution). This reaction produces HCHO with yields comparable to that of ISOPO₂ + NO and may be an important source in very-low NO_x regimes. Second, AM3 assumes a constant ISOPO₂ isomer distribution and thus underpredicts the isomerization rate relative to MCMv3.3.1, especially at mid to high NO_x (Fig. S7D). AM3 also includes HCHO and other small oxidized VOC as direct products of isomerization rather than producing hydroperoxyaldehydes and other large products, which influences the timescale of HCHO production and thus the partitioning between prompt and background HCHO. The impact of the RO₂ reaction and isomerization channels on HCHO yields is likely minor but depends significantly on the RO₂/HO₂ ratio (at low NO_x) and on the overall ISOPO₂ lifetime, which affects the ISOPO₂ isomer distribution. For the particular model conditions in Fig. S3B, ISOPO₂ lifetimes for the two mechanisms can differ by as much as 25% at the lowest NO_x values (Fig. S7E). Regardless of these differences, the results shown in Fig. 4 confirm that both the condensed AM3 and explicit MCMv3.3.1 mechanisms perform similarly with regard to overall HCHO production.

6 Mechanistic Drivers of the NO_x – HCHO Relationship

Despite the complexity of gas-phase organic chemistry, the impact of NO_x on HCHO production essentially reduces to two factors: radical cycling and RO₂ branching. Increasing NO enhances the conversion of HO₂ to OH (R4) and thus accelerates VOC oxidation (R5). RO₂ is also produced, to a lesser extent, by VOC ozonolysis and photolysis (R6). Subsequent production of HCHO depends on the structure and fate of RO₂ intermediates, which can react with NO, HO₂, other RO₂, or isomerize (R7).





Here, α represents a bulk branching ratio for HCHO production weighted over all RO_2 reactions. The RO_2 lifetime is typically less than 100 s during the day, so (R5) is the rate-limiting step in HCHO formation. The HCHO production rate is then equal to the product of the total RO_2 production rate and the bulk branching ratio:

$$P(\text{HCHO}) = \alpha P(\text{RO}_2) \quad (3)$$

Though total RO_2 losses include reactions that do not make HCHO, α is still a useful metric for the relationship between HCHO production and overall VOC oxidation.

To disentangle these factors, we extract chemical rates from the diel steady-state UWCM simulations discussed in Sect. 5. Figure 5A shows the gross production rates for total peroxy radicals and HCHO as a function of NO_x . Consistent with our earlier discussion, HCHO production increases by a factor of 3 from low to high NO_x . Total RO_2 production increases by a factor of 2 over this same range, driven primarily by increasing OH. The bulk branching ratio α , calculated as the ratio of HCHO and RO_2 production rates, increases from 0.43 to 0.62 (Fig. 5B). This trend is consistent with NO_x -dependent branching ratios of several major HCHO precursors, including isoprene hydroxyperoxy radicals (ISOPO_2) and methyl peroxy radical (Fig. 5B). Based on this analysis, we conclude that enhanced OH production is the main driver for the NO_x dependence of HCHO production, with variations in RO_2 branching playing a lesser (but still important) role.

Using a combination of regional modeling and satellite observations, a recent study by Valin et al. (2016) also examines the drivers of HCHO production. They concur that OH production exerts a controlling influence on HCHO throughout the Southeast U.S. In contrast to our study, however, they assert that changes in RO_2 branching have a negligible effect on the HCHO- NO_x dependence. There are several potential explanations for this discrepancy. First, Valin et al. (2016) derive an “effective branching ratio” that is analogous to the bulk branching ratio in Eqn. (3) but calculated with reference to production of OH rather than RO_2 . Many OH sinks do not form RO_2 radicals (e.g. reaction with CO, HCHO, methanol and NO_2) and thus will not make HCHO. The fractional contribution of such reactants to total modeled OH reactivity increases from 36% to 60% over our NO_x range; thus, using $P(\text{OH})$ instead of $P(\text{RO}_2)$ to calculate α from Eqn. (3) would effectively normalize out the NO_x dependence of RO_2 branching (Fig. 5B). Second, these two studies use very different photochemical

mechanisms. Valin et al. (2016) use a modified version of the lumped Regional Atmospheric Chemistry Mechanism 2 (RACM2) (Browne et al., 2014; Goliff et al., 2013), while our box model uses the explicit MCMv3.3.1 (Jenkin et al., 2015). In Valin et al. (2016), it is stated that increasing HCHO production from the $\text{RO}_2 + \text{RO}_2$ channel compensates for decreasing production from $\text{RO}_2 + \text{NO}$ – an effect that we do not observe. Deeper investigation reveals that the rate constant for reaction of ISOPO_2 with HO_2 in RACM2 is a factor of 2 lower than those used in both MCMv3.3.1 and the AM3 mechanism, which are based on the experimentally-derived parameterization of Boyd et al. (2003). Thus, our model predicts a significantly larger contribution of $\text{RO}_2 + \text{HO}_2$ (which produces negligible HCHO) to the total RO_2 sink. These differences highlight the importance of carefully evaluating chemical mechanisms before using models to interpret *in situ* and satellite observations.

Increased OH also reduces the lifetime of HCHO, which may affect the HCHO budget if this reaction becomes competitive with photolysis. UWCM predicts an average HCHO photolysis lifetime of 4 hours and OH reaction lifetimes that range from 3 hours at high NO_x to 12 hours at low NO_x . Thus, photolysis is typically the dominant loss process and the scaling of HCHO lifetime with OH is typically weak. The net chemical tendency of HCHO (production minus loss) is positive and increasing throughout the range of model NO_x conditions. Faster loss due to reaction with OH therefore only slightly dampens the enhancement in HCHO production.

7 Conclusions

Using SENEX aircraft observations, we have quantified the NO_x dependence of the relationship between isoprene emission strength and HCHO mixing ratios. Simultaneous measurements of isoprene, MVK and MACR define a photochemical clock for isoprene oxidation, allowing separation of prompt HCHO production (which retains the isoprene source signature) and background HCHO from late-generation isoprene oxidation products, methane and other long-lived VOC. The prompt HCHO yield increases by a factor of 3 (0.3 to 0.9 ppbv ppbv⁻¹) and the average background HCHO mixing ratio doubles (1.6 to 3.3 ppbv) over the range of NO_x values encountered in the southeast U.S. (0.1 – 2 ppbv). This analytical method is applied to evaluate the performance of a global chemical transport model and a 0-D steady-state box model. Both models accurately reproduce the observed NO_x trend of the prompt HCHO yield, indicating that both chemical mechanisms accurately capture early-stage

1 isoprene oxidation. On the other hand, both models also under-predict background HCHO
2 abundance by 0.5 – 1 ppbv, which is a significant fraction of total HCHO in some cases. This
3 suggests insufficient build-up of isoprene-derived long-lived precursors in the models,
4 missing VOC not related to isoprene, or insufficient OH. Box model results also provide
5 insight into the mechanistic drivers of the observed NO_x trends. Over the NO_x range studied
6 here, a 100% increase in total RO₂ production and a 40% increase in the HCHO production
7 branching ratio give rise to a 3-fold increase in total HCHO production.

8 To our knowledge, there are no direct laboratory measurements of HCHO yields from
9 low-NO_x isoprene chemistry; thus, the results presented here constitute the first measurement-
10 constrained evaluation of the isoprene-HCHO link across NO_x regimes. The AM3 and
11 MCMv3.3.1 mechanisms differ substantially (the former is highly condensed while the latter
12 is explicit), but both contain recent updates to isoprene degradation. We expect that other
13 mechanisms will also perform well if they accurately reflect our current best understanding.
14 The observations presented here do not include the extremely-low NO_x regime (NO_x < 0.1
15 ppbv) typical of remote regions like the Amazon and equatorial Africa. In such pristine
16 regions, smearing of HCHO production is expected to be more severe (Barkley et al., 2013),
17 and total HCHO production may be significantly lower if the RO₂ fate favors
18 functionalization over fragmentation (e.g. isomerization). More work is needed to map out
19 this area of the urban-rural spectrum. It may also be possible to apply the methods developed
20 here to evaluate the chemistry of glyoxal, another key tracer of VOC oxidation that is also
21 amenable to orbital observations (Kaiser et al., 2015b; Li et al., 2016) and is believed to be an
22 important precursor for SOA (McNeill et al., 2012).

23 These results also carry implications for top-down isoprene emission estimates.
24 Uncertainties in low-NO_x chemistry are often cited as the largest source of potential error in
25 derived emissions (Marais et al., 2012; Palmer et al., 2006). Based on our analysis, current
26 mechanisms appear to capture low-NO_x production of HCHO, MVK and MACR, thus such
27 errors are likely less severe than commonly asserted. Recent work has acknowledged the
28 impact of NO_x on the prompt yield of HCHO from isoprene (Marais et al., 2012). We
29 advocate considering the NO_x dependence of background HCHO as well, since this can
30 constitute a significant fraction of the total HCHO column. For scale, the derived background
31 HCHO mixing ratio of 1.6 – 3.3 ppbv is 37 – 77% of the campaign-mean observed HCHO
32 mixing ratio of 4.3 ppbv. Forthcoming geostationary observations will resolve local gradients

1 in chemical regime, and smearing and background HCHO production will become
2 problematic even in high-NO_x regions. Indeed, even current-generation orbital instruments are
3 capable of resolving some urban-rural gradients in HCHO columns (Boeke et al., 2011).
4 When applying advanced statistical techniques like inversion, model results will only be as
5 accurate as the chemical mechanisms driving them. Continued field observations are crucial
6 for providing confidence in our ability to link HCHO to its sources. In this regard, recent
7 work has highlighted the potential of airborne eddy covariance fluxes to quantify both
8 surface-atmosphere exchange and *in situ* chemical processes (Karl et al., 2013; Kaser et al.,
9 2015; Misztal et al., 2014; Wolfe et al., 2015). With such tools, it should be possible to
10 simultaneously measure both isoprene emissions and HCHO columns, thereby obtaining a
11 direct experimental constraint on the link between these two quantities.

13 **Data Availability**

14 All data used in this study are publicly accessible on the SENEX website
15 (<http://www.esrl.noaa.gov/csd/projects/senex/>).

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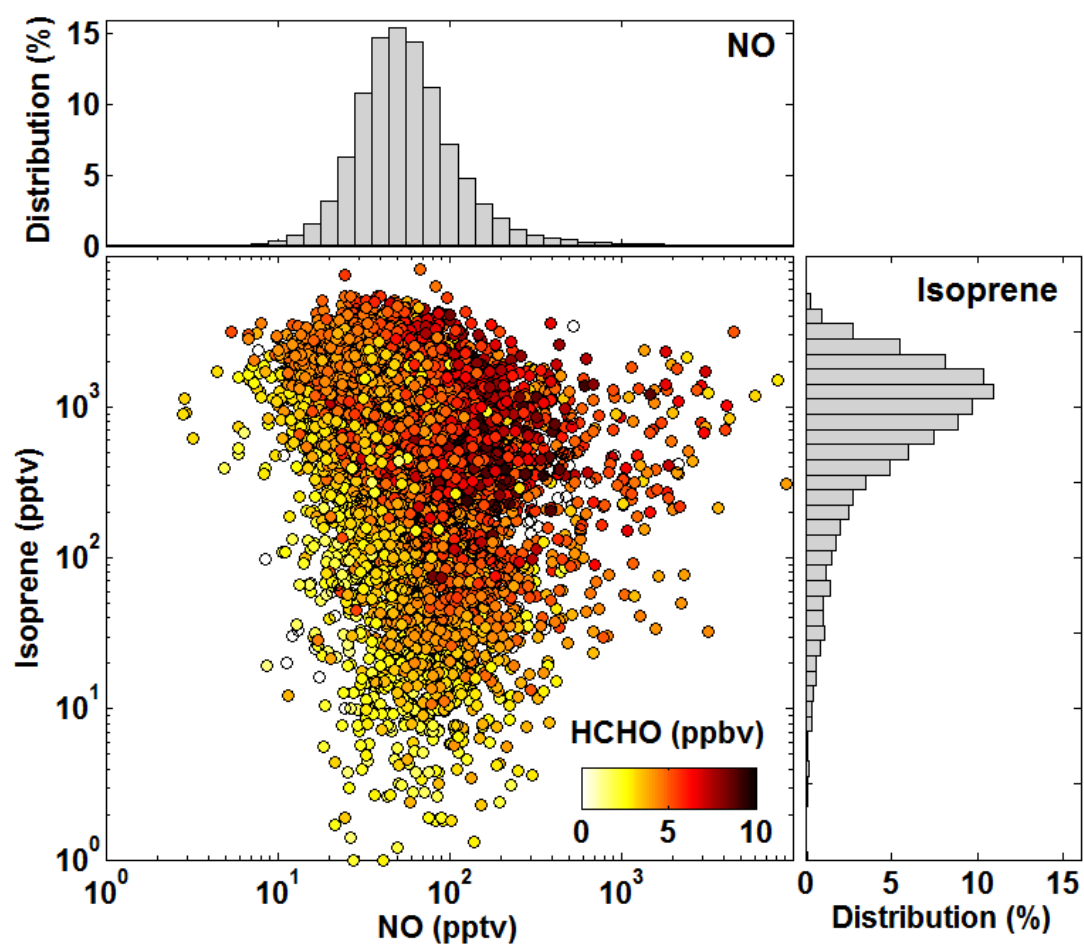


Figure 1. Co-variation of isoprene, NO and HCHO mixing ratios in the summertime Southeast U.S. Data are limited to daytime boundary layer observations. Histograms show the corresponding NO and isoprene distributions.

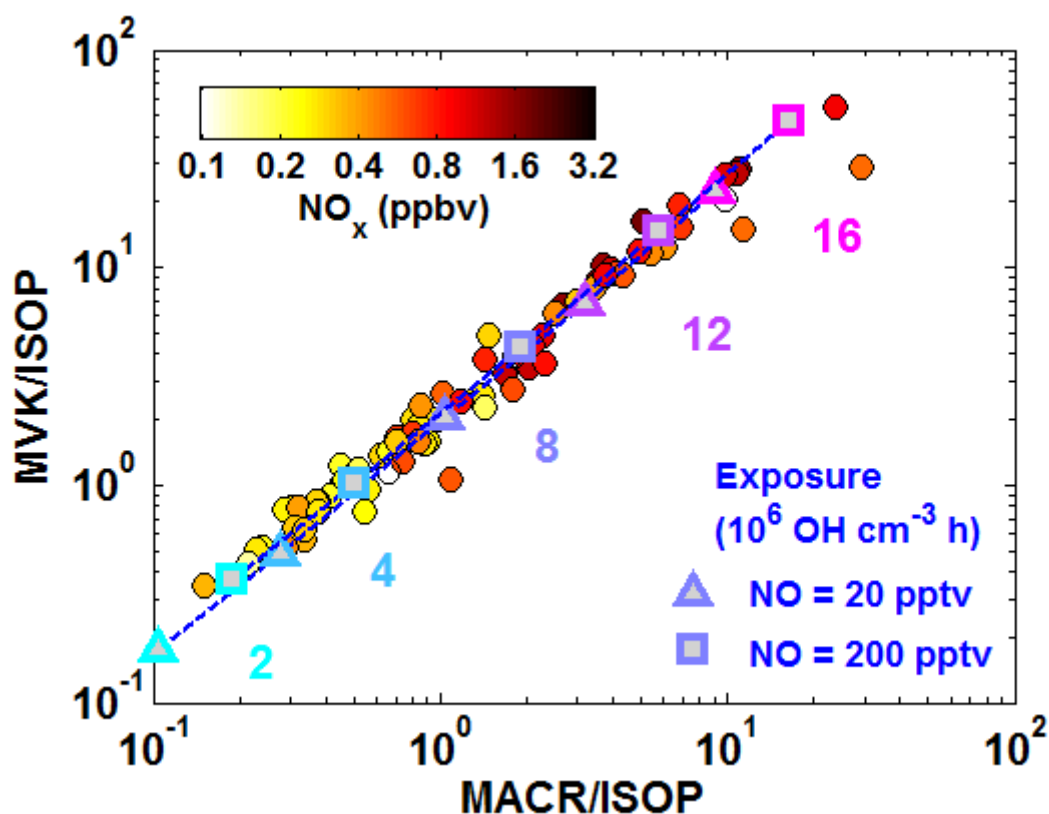


Figure 2. A photochemical clock of isoprene oxidation defined by the progression of daughter/parent ratios. Solid circles show the observed ratios calculated from iWAS observations, colored by NO_x . Blue/purple symbols, dashed lines, and text indicate the theoretical exposures (the product of OH concentration and time) corresponding to any given daughter/parent relationship. Theoretical values are calculated at 298K using MVK and MACR yields for NO values of 20 pptv (triangles) and 200 pptv (squares).

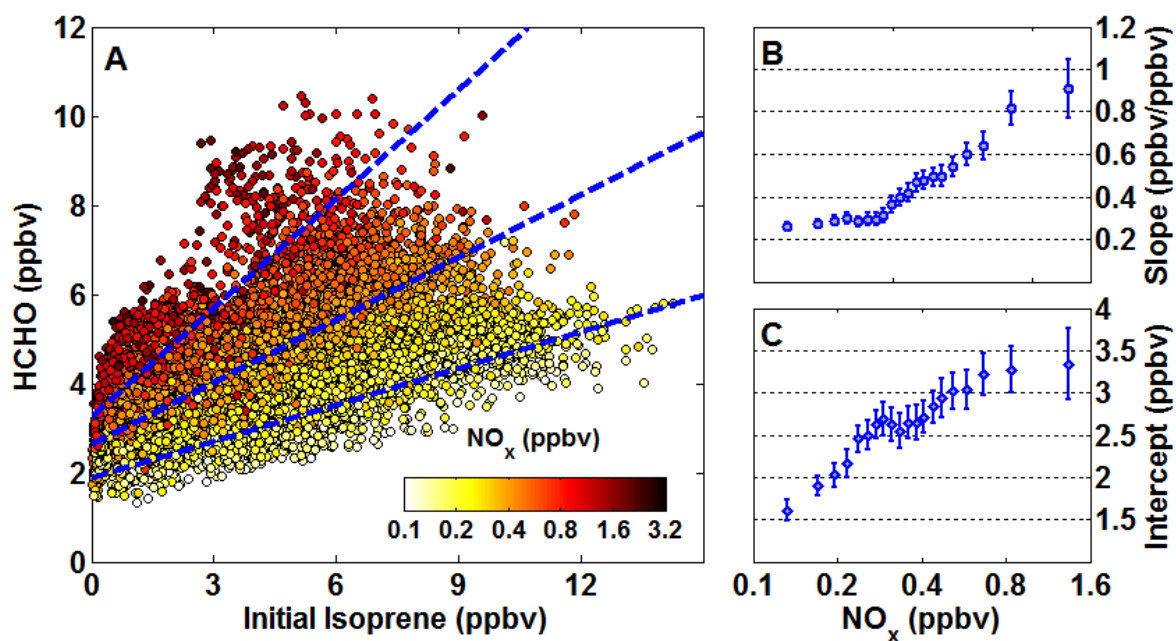


Figure 3. (A) NO_x modulates the relationship between observed HCHO and calculated initial isoprene mixing ratios. Symbols denote all 1-second data. Dashed lines illustrate representative major-axis fits of NO_x-grouped subsets at mean NO_x values of 170, 380 and 810 pptv (see text for details of fitting procedure). The slope (B) and intercept (C) of these fits are the prompt HCHO yield and background HCHO mixing ratio, respectively. Error bars in (B) and (C) are 3σ fitting uncertainties.

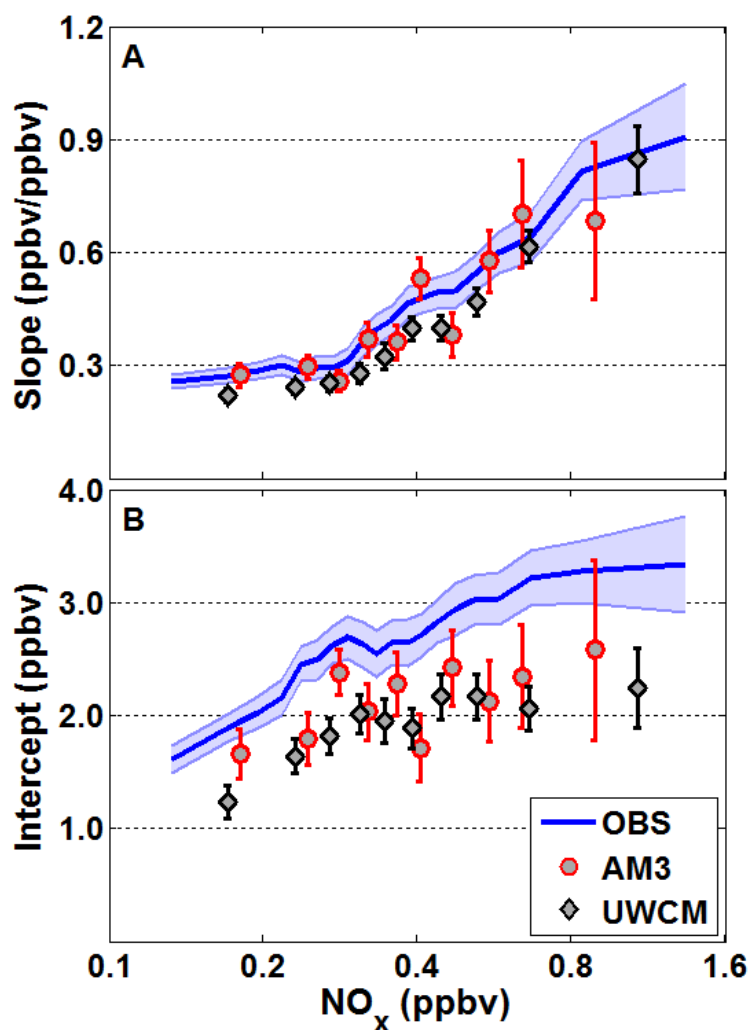
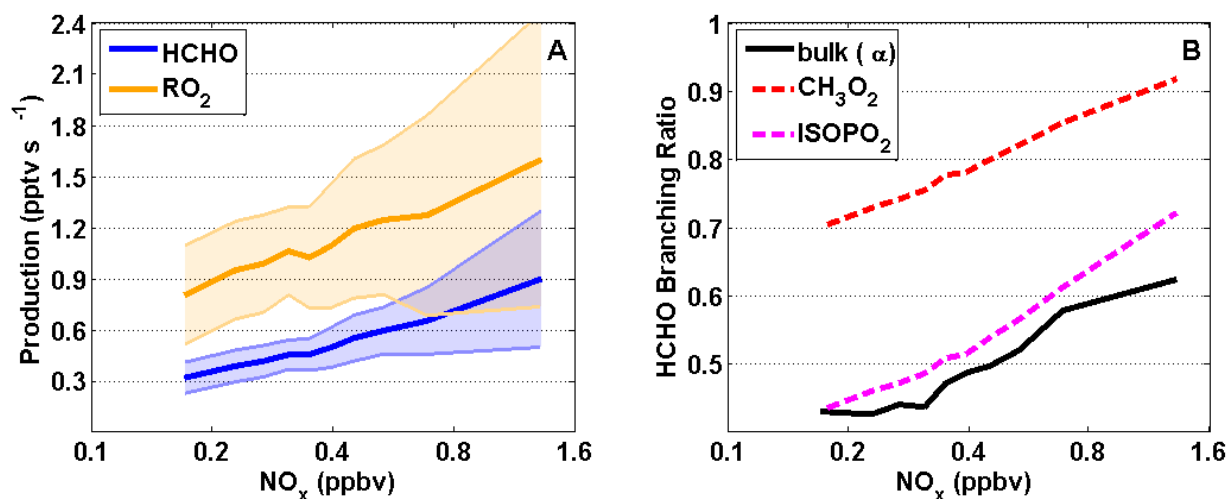


Figure 4. Comparison of observed and model-derived relationships between HCHO and initial isoprene versus NO_x . Slopes (A) and intercepts (B) are calculated as described in the text. The observed values (blue line with shading) are the same as those shown in Figs. 3B-C. Symbols represent fit results for the global AM3 model (red circles) and the 0-D UWCM box model (black diamonds). Error bars denote 3 σ fitting uncertainties.

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4 **Figure 5.** NO_x dependence of chemical properties related to HCHO production, extracted
 5 from the UWCM simulation of SENEX observations. (A) Production rates for HCHO (blue)
 6 and total RO_2 (orange). (B) Branching ratios for HCHO production weighted over all RO_2
 7 (solid black line) and for several individual RO_2 , including methyl peroxy radical (red) and
 8 total isoprene hydroperoxy radicals (magenta). All quantities are averaged over NO_x using
 9 10 bins with equal numbers of points. In (A), solid lines show the mean and shading is 1 σ
 10 variability.